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Figure 12.1.1 illustrates this procedure for a solution of cobalt (II) chloride dihydrate in

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ethanol. Note that the volume of the solvent is not specified.

Because the solute occupies space in the solution, the volume of the solvent needed is almost always less than the desired volume of solution.

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#### SHORT ANSWER

Answer the following questions in the space provided.

1. Match the type of mixture on the left to its representative particle diameter on the right.

c solutions	(a) larger than 1000 nm
a suspensions	(b) 1 nm to 1000 nm
b colloids	(c) smaller than 1 nm

2. Identify the solvent in each of the following examples:

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### SECTION 1 SHORT

ANSWER 1. c 2. a 3. b  
2. a. alcohol b. water c.  
the gels 3. The mixture  
is a colloid. The  
properties are  
consistent with those  
reported in Table 3 on  
page 404 of the text.  
The particle size is  
small, but not too  
small, and the mixture

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Solutions Worksheet is suitable for 9th - 12th Grade. Here is a different approach to solutions: a worksheet that has chemistry learners use words to describe them! This is an outstanding exercise that will stretch learners to show what they know. They answer questions comparing solutions, suspensions, and colloids.

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CHM 151: Chapter 12  
Notes page 1 of 9

Chapter 12: Solutions  
Problems: 3, 5, 8, 12,  
14, 16, 22, 29, 32,  
41-58, 61-68, 71-74

solution: homogeneous  
mixture of a solute  
dissolved in a solvent  
solute: component  
present in smaller  
amount solvent:  
component present in

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greater amount

**Solubility:** Maximum

amount of solute

dissolved in solvent at  
specific temp.

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CHAPTER 12

SOLUTIONS MANUAL

States of Matter

Section 12.1 Gases

pages 402–410

Practice Problems page

405 1. Calculate the

ratio of effusion rates

for nitrogen ( $N_2$ ) and

neon (Ne).  $\text{Rate } N / \text{Rate}$

Ne  $0.849 \text{ Rate } N \_ \text{Rate}$

Ne  $\sqrt{\text{molar mass Ne}} \_ \_$

$\text{molar mass N} \sqrt{\_} \_ 20.18$

$28.02 \cdot 0.849 \cdot 2.$

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### Section 1

Calculate the ratio of diffusion rates for

carbon monoxide and carbon dioxide. 1.25

Rate CO \_ Rate CO 2  $\sqrt$

### **States of Matter**

Section 1- Types of Mixtures Solutions are homogeneous mixtures of two or more substances in a single phase. Soluble describes a substance as capable of being dissolved. Solvent is the dissolving medium

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## Chapter 12

### Section 1

in a solution. Solute is the substance that is dissolved in a solution.

## **Modern Chemistry**

### **Chapter 12-**

### **Solutions**

Chemistry Chapter 12  
Section 1 and 2.

STUDY. PLAY. A

homogeneous mixture of two or more substances in a single phase. solution. The dissolving medium in a solution or the substance present in a



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## Section 1

larger amount. solvent.

The substance

dissolved in a solution

or the substance

present in the smaller

amount. solute.

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... Continue the backtracking search for a solution to the four-queens problem, which was started in this section, to find the second solution to the problem. b. Explain how the board's symmetry can be used to find the second solution to the ...

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**Electricity ...**

statistics and  
probability. statistics  
and probability  
questions and answers.

Reserve Problems

Chapter 12 Section 1

Problem 11 An Article

In Technometrics

(1974, Vol. 16, ...

Question: Reserve

Problems Chapter 12

Section 1 Problem 11

An Article In

Technometrics (1974,

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Section 1  
Vol. 16, Pp. 523-531)  
Solutions  
Considered The  
Following Stack-loss  
Data From A Plant  
Oxidizing Ammonia To  
Nitric Acid.

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Create. Chapter 12  
Section 1,2,3. ... what  
is the concentration of  
100. mL aqueous  
solution that contains  
1.00 g KCl (molar  
mass= 74.55g/mol)  
0.134M. What is the  
molarity of a solution  
composed of 8.21- g of  
...

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Exercise 12.3 and  
Exercise 12.4 in  
English Medium and  
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Prashnavali 12.2. 2,  
Prashnavali 12.3 and  
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Section 1

CHAPTER 12 . REVIEW

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ANSWER Answer the following questions in the space provided. 1.

Describe the errors made by the following students in making molar solutions. 3.

James needs a 0.600 M solution of KCl. He measures out 0.600 g of KCl and adds 1 L of water to the solid.

James made several errors.

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## Chapter 12

### Section 1

## Solutions

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